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Stoichiometry: Limiting Reactant

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Reagents Practice #2-6 Answers
PRACTICE: Stoichiometry ...

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Limiting reagents only (two given reactants, one wanted product) Mix & match (both simple stoichiometry and

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limiting reagent problems) Units to use

(select at least one): Grams. Moles.

Particles (e.g.

atoms/molecules/formula units)

Chemical formulas or names:

Formulas only. Names only.

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c) What is the limiting reagent in
problem #2? d) How much of the
excess reagent will be left over after

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the reaction is complete? Find the molarity of the following solutions:
1) 0.5 moles of sodium chloride is dissolved to make 0.05 liters of solution.
2) 0.5 grams of sodium chloride is dissolved to make 0.05 liters of solution.

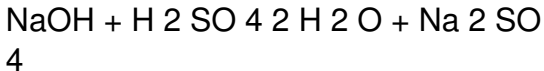
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Stoichiometry Practice Worksheet

Balancing Equations and Simple
Stoichiometry-KEY Balance the
following equations: 1) $1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3$
2) $2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O}$
3) $1 \text{ HBr} + 1 \text{ KHCO}_3 \rightarrow 1 \text{ H}_2\text{O} + 1 \text{ KBr} + 1 \text{ CO}_2$
4) $2 \text{ GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 \rightarrow 1 \text{ Ga}_2(\text{SO}_3)_3 + 6 \text{ NaBr}$

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3 Using the following equation: 2



~~Balancing Equations and Simple
Stoichiometry KEY~~

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will many period as we run by before.
You can do it even though work
something else at home and even in
your workplace. consequently easy!
So, are you question? Just exercise
just what we meet the expense of
below as skillfully as evaluation

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9) What is the limiting reagent for
problem 6? _____ 10) How much of
the excess reagent is left over after the

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Reaction from problem 6 is finished?

11) If 35 grams of carbon dioxide are actually formed from the reaction in problem 6, what is the percent yield of this reaction?

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In order to determine the limiting

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Reactants, we need to determine which of the reactants will give less product. According to the balanced chemical equation, every 2 moles of H_2 will yield 2 moles of H_2O . Remember, this is determined based on the mole ratio of H_2 and H_2O , which is 2:2 (the coefficients) in front of each molecule.

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1) For the reaction $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2$

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H₂O(g), how many liters of water can be made from 5 L of oxygen gas and an excess of hydrogen? 10 L. 5L O₂ x 2L H₂O/1L O₂ =. Gas Stoichiometry Practice Sheet Stoichiometry Practice Answer Key Chemfiesta PDF Download ...

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analysis: Volatilization gravimetry
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1) Determine limiting reagent: $\text{NBr}_3 \square$
50 "moles" / 2 = 25. $\text{NaOH} \square$ 57

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"moles" / 3 = 19. NaOH is the limiting reagent. Note that there need be no conversion from grams to moles. Discussions of numbers of molecules uses numbers that are directly proportional to the number of moles and do not need to be converted.

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